

Objective: Students understand that compounds are fixed ratios of elements bonded together. Students know how to read a chemical formula

Do Now: You have a test tube of H_2O and a test tube of H_2O_2 . Would you classify these as compounds or mixtures? What do these formulas tell us?

1. Has a variable composition

- A. Mixture
- B. Element
- C. Compound

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2. Made up of a fixed ratio of bonded elements

- A. Mixture
- B. Element
- C. Compound

3. Cannot be divided by physical means

- A. Mixture
- B. Element
- C. Compound
- D. Both mixtures and compounds
- E. Both Mixtures and compounds

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4. Cannot be separated by chemical means

- A. Mixture
- ~~B. Element~~
- C. Compound
- D. Both compounds and elements
- E. Both Mixtures and compounds

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5. Which box below shows a single type of compound?

- ~~A. only A~~
- B. only B
- C. only C
- D. only D
- E. A and B
- F. B and D

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6. Which boxes below represent an element?


- A. only A
- B. only B
- C. only C
- D. only D
- E. A and B
- ~~F. B and D~~

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7. Which box represents a mixture?

- A. only A
- B. only B
- ~~C. only C~~
- D. only D
- E. A and B
- F. B and D

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Dalton's Atomic Theory 

- 1) All matter is made of atoms.
- 2) All atoms of a given element are identical in properties
- 3) Compounds are formed by the joining of two or more different kinds of atoms
- 4) When elements react to form compounds, they react in defined, whole number ratios

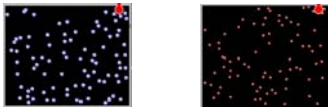
$H_2O \rightarrow \begin{matrix} H & O \\ H & \end{matrix}$
 $2 : 1$

$2 H_2O \rightarrow \begin{matrix} H & H & O & O \\ H & H & & \end{matrix}$
 $2 : 1$

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Avogadro's Hypothesis

1. What is reasonable to conclude about the number of gas particles in each container if the pressure and temperature is the same in both containers?



$P = , V = , T = , \# \text{ particles} =$

\Rightarrow \Rightarrow

If $P = , V = , T =$ then $\# \text{ particles} =$

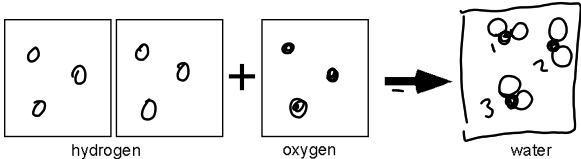
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Avogadro reached this same conclusion building on the work of Gay-Lussac, who first noted that gases (at the same T and P) reacted in simple integer volume ratios. His hypothesis made it possible to deduce the formulas of compounds formed when these gases react.

You have seen evidence that two volumes of hydrogen gas react with one volume of oxygen gas (at the same T and P) to produce water. The conclusion that two *molecules* of hydrogen combine with one *molecule* of oxygen to form water works only if we assume that each volume of gas contains the same number of particles.

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2. Represent molecules of hydrogen and oxygen in the containers below. React these molecules to form water molecules, leaving no leftover gas.



hydrogen + oxygen = water

H_2O

What do the H, O, and 2 in the chemical formula tell us about the composition of water?

H_2O
 \swarrow \downarrow \searrow
 hydrogen # of atoms oxygen

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3. In like manner, represent particle diagrams that account for the fact that one volume of hydrogen combines with one volume of chlorine to form hydrogen chloride. What do you suppose is the formula of hydrogen chloride?

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4. Represent the reaction in which one volume of nitrogen gas reacts with three volumes of hydrogen gas to form ammonia. What is the formula for ammonia?

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Part 2
Chemists occasionally found that one volume of gas A reacted with one volume of gas B to produce *two volumes* of gaseous product. Early chemists like Gay-Lussac, were unable to account for this behavior of gases. Avogadro's key contribution was that he reasoned that the molecules of some gaseous elements *must contain two atoms*.

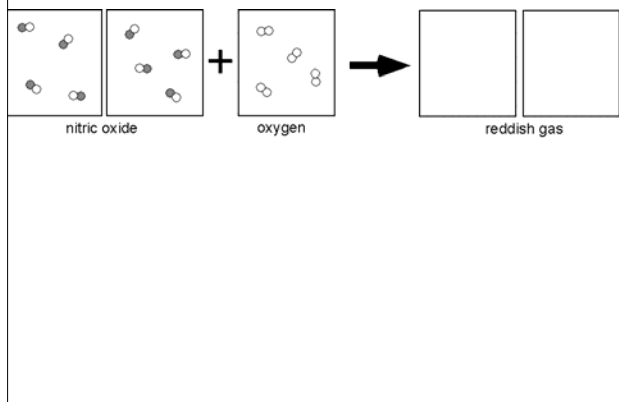
5. Consider the reaction between hydrogen and chlorine. *Two* volumes of hydrogen chloride are formed. Sketch particle diagrams consistent with Avogadro's Hypothesis to represent this reaction. Explain why hydrogen and chlorine molecules that have only one atom each cannot account for the observed behavior.

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6. When two volumes of hydrogen gas react with one volume of oxygen gas, two volumes of gaseous water are formed. Modify the diagram you made for #2 to represent molecules of hydrogen, oxygen and water in this reaction.

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7. Two volumes of nitric oxide react with one volume of oxygen gas to form two volumes of a reddish-brown gas. Deduce the formula of this gas and sketch particle representations of its molecules.



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