

Get out ion sheet and homework.

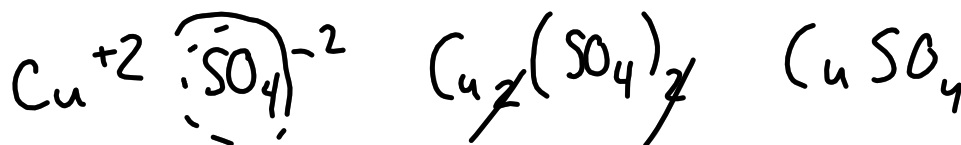
Fix last exit ticket:

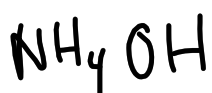
Write the name for the following formulas:

1. NaF sodium fluoride
- ② CuO copper (II) oxide -2
- ③ LiCO₃ lithium carbonate

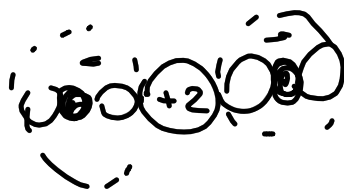
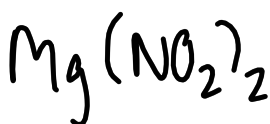
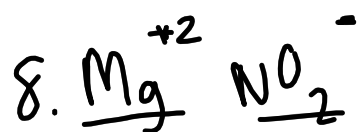
Write the formula for the following compounds:

- ④ Calcium Chloride $\text{Ca}^{+2} \text{Cl}^{-}$ CaCl_2
- ⑤ Tin (IV) Oxide $\text{Sn}^{+4} \text{O}^{-2}$ Sn_2O_4 SnO_2
- ⑥ Silver Nitrate $\text{Ag}^{+} \text{NO}_3^{-}$ AgNO_3
- ⑦ Copper (II) Sulfate





ammonium hydroxide



magnesium nitrite

How to Write a chemical equation:

- Reactant 1 + Reactant 2 -----> Product
- Reactant -----> Product 1 + Product 2
- Reactant 1 + Reactant 2 -----> Product 1 + Product 2

Reactants → what you start with in a
chemical reaction

→ indicates a change

product → what you end with

Symbols in chemical equations

(s) = solid

(l) = liquid

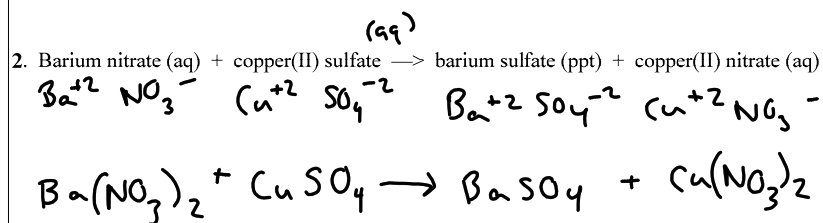
(g) = gas

(aq) = aqueous → dissolved in water (solution)

(ppt) = precipitate → solid that forms when
you mix two aq
solutions

Write the chemical formula for each compound in the reaction.
Under the chemical formula for the precipitate (ppt), write the **color** of the precipitate.

1. Nickel(II) chloride(aq) + sodium sulfide(aq) \rightarrow nickel(II) sulfide (ppt) + sodium chloride (aq)



3. Sodium carbonate(aq) + calcium chloride(aq) \rightarrow calcium carbonate (ppt) + sodium chloride (aq)

4. Potassium chromate (aq) + silver nitrate (aq) \rightarrow silver chromate (ppt) + potassium nitrate (aq)

5. Silver nitrate (aq) + nickel(II) chloride (aq) \rightarrow silver chloride (ppt) + nickel nitrate (aq)

6. Cobalt(II) nitrate and sodium hydroxide (aq) \rightarrow cobalt(II) hydroxide (ppt) + sodium nitrate (aq)

7. Potassium iodide (aq) and lead(II) nitrate (aq) \rightarrow lead(II) iodide (ppt) + potassium nitrate (aq)